

=> FILE WPIX

FILE 'WPIX' ENTERED AT 11:55:43 ON 20 SEP 2004  
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FILE LAST UPDATED: 15 SEP 2004 <20040915/UP>  
MOST RECENT DERWENT UPDATE: 200459 <200459/DW>  
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=> D QUE L49

L22 1 SEA FILE=REGISTRY ABB=ON 12634-54-7  
L25 1 SEA FILE=REGISTRY ABB=ON MAGNESIUM/CN  
L34 1543680 SEA FILE=HCAPLUS ABB=ON L22 OR L25 OR (MG OR MAGNESIUM)  
L39 6125 SEA FILE=WPIX ABB=ON L34(6A) (MELT? OR MOLTEN? OR LIQ?)  
L42 3 SEA FILE=WPIX ABB=ON L39 AND (DIFLUOROMETHANE OR PENTAFLUOROET  
HANE OR DIFLUOROETHANE OR HEPTAFLUOROPROPANE OR (METHOXY OR  
ETHOXY) (W)NONAFLUOROBUTANE OR DIHYDRODECAFLUOROPENTANE OR  
TETRAFLUOROETHANE)  
L43 0 SEA FILE=WPIX ABB=ON L39 AND (HYDROFLUOROETHER# OR HYDRO(W)FLU  
ORO(W)ETHER# OR HYDROFLUORO(W)ETHER#)  
L44 3 SEA FILE=WPIX ABB=ON L42 OR L43  
L46 1 SEA FILE=COMPENDEX ABB=ON L39 AND HYDROFLUORO?  
L49 4 SEA FILE=WPIX ABB=ON L44 OR L46

=> D L49 FULL 1-4

L49 ANSWER 1 OF 4 WPIX COPYRIGHT 2004 THOMSON DERWENT on STN  
AN 2004-227174 [21] WPIX  
DNN N2004-179528 DNC C2004-089559  
TI Production of **magnesium** alloy strip comprises passing  
**molten** alloy to feeding device, feeding molten alloy to chamber,  
rotating rolls in which alloy is drawn from chamber, and flowing coolant  
fluid through each roll.  
DC M22 P53  
IN ALLEN, R V; BORBRIDGE, W; EAST, D R; LIANG, D D  
PA (CSIR) COMMONWEALTH SCI & IND RES ORG  
CYC 105  
PI WO 2004020126 A1 20040311 (200421)\* EN 33 B22D011-06  
RW: AT BE BG CH CY CZ DE DK EA EE ES FI FR GB GH GM GR HU IE IT KE LS  
LU MC MW MZ NL OA PT RO SD SE SI SK SL SZ TR TZ UG ZM ZW

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DM DZ EC EE ES FI GB GD GE GH GM HR HU ID IL IN IS JP KE KG KP KR  
KZ LC LK LR LS LT LU LV MA MD MG MK MN MW MX MZ NI NO NZ OM PG PH  
PL PT RO RU SC SD SE SG SK SL SY TJ TM TN TR TT TZ UA UG US UZ VC  
VN YU ZA ZM ZW

ADT WO 2004020126 A1 WO 2003-AU1097 20030827

PRAI AU 2002-951075 20020829

IC ICM B22D011-06

AB WO2004020126 A UPAB: 20040326

NOVELTY - **Magnesium** alloy strip is produced by passing **molten** alloy from source of supply to feeding device, feeding molten alloy from feeding device through a nozzle (30) to a chamber (44), rotating rolls (32ab) in opposite direction in which alloy is drawn from the chamber, and flowing coolant fluid through each roll to provide internal cooling of the rolls.

DETAILED DESCRIPTION - Production of **magnesium** alloy strip involves passing **molten** alloy from source of supply to feeding device; feeding the molten alloy from the feeding device through the nozzle to the chamber formed between an elongate outlet of the nozzle and a pair of parallel rolls that are spaced one above the other to define a bite; rotating the rolls in opposite direction in which alloy is drawn from the chamber through the bite; and flowing coolant fluid through each roll to provide internal cooling of the rolls to cool the alloy received in the chamber by heat energy extraction by the cooled rolls to complete solidification of the alloy. The production method further includes maintaining the alloy held at the source at a temperature to maintain the alloy in the feed device at superheated temperature above its liquidus temperature, maintaining a depth of molten alloy in the feed device at a controlled constant height of 5-22 mm above a centerline of the bite in a plane containing the axes of the rolls, and maintaining heat energy extraction by the cooled rolls at a level to maintain alloy strip issuing from the bite at a surface temperature below 400 deg. C.

USE - Used in the production of magnesium alloy strip.

ADVANTAGE - Provides magnesium alloy strip of a required thickness and width.

DESCRIPTION OF DRAWING(S) - The figure shows an enlarged scale details relating to magnesium alloy solidification.

Nozzle 30

Rolls 32ab

Nip 34

Horizontal plates 36,37

Outlet 42

Chamber 44

Dwg.9/12

TECH WO 2004020126 A1UPTX: 20040326

TECHNOLOGY FOCUS - METALLURGY - Preferred Methods: The alloy at the source is held at 15-60 degrees C above its liquidus temperature. The level of heat extraction is able to maintain the surface temperature at 180-300 degrees C. The rolls apply a specific load of 2-500 kg/mm of roll length to the solidified alloy passing through the bite. A protective atmosphere is maintained over the molten alloy to safeguard against oxidation and risk of fire. The atmosphere includes a minor proportion of 1,1,1,2-tetrafluoroethane.

FS CPI GMPI

FA AB; GI

MC CPI: M22-G03A1A

L49 ANSWER 2 OF 4 WPIX COPYRIGHT 2004 THOMSON DERWENT on STN

AN 2002-055485 [07] WPIX

DNN N2002-040877. DNC C2002-015901...

TI Treatment of **molten magnesium** to protect **magnesium** from reacting with air involves exposing **molten magnesium** to gaseous mixture comprising fluorocarbon that includes limited amounts of hydrogen.

DC E16 M25 P53

IN MILBRATH, D S; OWENS, J G

PA (MINN) 3M INNOVATIVE PROPERTIES CO

CYC 96

PI WO 2001083836 A2 20011108 (200207)\* EN 26 C22B026-22

RW: AT BE CH CY DE DK EA ES FI FR GB GH GM GR IE IT KE LS LU MC MW MZ  
NL OA PT SD SE SL SZ TR TZ UG ZW

W: AE AG AL AM AT AU AZ BA BB BG BR BY BZ CA CH CN CO CR CU CZ DE DK  
DM DZ EE ES FI GB GD GE GH GM HR HU ID IL IN IS JP KE KG KP KR KZ  
LC LK LR LS LT LU LV MA MD MG MK MN MW MX MZ NO NZ PL PT RO RU SD  
SE SG SI SK SL TJ TM TR TT TZ UA UG UZ VN YU ZA ZW

AU 2001052990 A 20011112 (200222)

NO 2002000027 A 20020207 (200227) C22B000-00

KR 2002065466 A 20020813 (200309) C22B026-22

EP 1278897 A2 20030129 (200310) EN C22B026-22

R: AL AT BE CH CY DE DK ES FI FR GB GR IE IT LI LT LU LV MC MK NL PT  
RO SE SI TR

US 2003034094 A1 20030220 (200316) C23C014-12

US 6537346 B2 20030325 (200325) C22B009-05

US 2003164068 A1 20030904 (200359) A62D001-02

US 2003164069 A1 20030904 (200359) C22B009-00

MX 2001013455 A1 20020801 (200367) C22B026-22

JP 2003531965 W 20031028 (200373) 37 C22B026-22

US 6685764 B2 20040203 (200413) C22B009-05

US 6780220 B2 20040824 (200457) C22B026-22

ADT WO 2001083836 A2 WO 2001-US9899 20010328; AU 2001052990 A AU 2001-52990  
20010328; NO 2002000027 A WO 2001-US9899 20010328, NO 2002-27 20020103; KR  
2002065466 A KR 2002-700054 20020103; EP 1278897 A2 EP 2001-926456  
20010328, WO 2001-US9899 20010328; US 2003034094 A1 Provisional US  
2000-202169P 20000504, US 2001-780256 20010209; US 6537346 B2 Provisional  
US 2000-202169P 20000504, US 2001-780256 20010209; US 2003164068 A1  
Provisional US 2000-202169P 20000504, CIP of US 2001-780256 20010209, US  
2003-394852 20030322; US 2003164069 A1 Provisional US 2000-202169P  
20000504, CIP of US 2001-780256 20010209, US 2003-394853 20030322; MX  
2001013455 A1 WO 2001-US9899 20010328, MX 2001-13455 20011219; JP  
2003531965 W JP 2001-580443 20010328, WO 2001-US9899 20010328; US 6685764  
B2 Provisional US 2000-202169P 20000504, CIP of US 2001-780256 20010209,  
US 2003-394852 20030322; US 6780220 B2 Provisional US 2000-202169P  
20000504, CIP of US 2001-780256 20010209, US 2003-394853 20030322

FDT AU 2001052990 A Based on WO 2001083836; EP 1278897 A2 Based on WO  
2001083836; US 2003164068 A1 CIP of US 6537346; US 2003164069 A1 CIP of US  
6537346; MX 2001013455 A1 Based on WO 2001083836; JP 2003531965 W Based on  
WO 2001083836; US 6685764 B2 CIP of US 6537346; US 6780220 B2 CIP of US  
6537346

PRAI US 2001-780256 20010209; US 2000-202169P 20000504;  
US 2003-394852 20030322; US 2003-394853 20030322

IC ICM A62D001-02; C22B000-00; C22B009-00; C22B009-05; C22B026-22;  
C23C014-12

ICS B22D001-00; B22D021-02; B22D021-04; C07C049-16; C22B009-02;  
C23C008-06; C23C026-00; G06F017-60

AB WO 200183836 A UPAB: 20020130

NOVELTY - **Molten magnesium** is treated by providing **molten magnesium** and exposing the **magnesium** to a gaseous mixture comprising fluorocarbon such as perfluoroketone, and/or

**hydrofluoroketone.**

DETAILED DESCRIPTION - INDEPENDENT CLAIMS are also included for:

(A) solid magnesium including a film formed on the surface of the magnesium including a reaction product of molten magnesium and a fluorocarbon; and

(B) a method of extinguishing fire on the surface of magnesium.

USE - For protecting magnesium used in automotive and aerospace industries, from reacting with oxygen in air.

ADVANTAGE - The global warming potentials of PFKs and HFKs are quite low (less than 22200). Thus, the process is more environmentally friendly. PFKs and HFKs also react more fully with molten

magnesium than does sulfur hexafluoride. AS a result less unreacted cover gas can be emitted to the atmosphere and/or less cover gas can be required to produce a comparably performing protective film. As a result, useful concentrations of the cover gas can be lowered, reducing the global warming impact. The full substitution of the invention for SF6 can be accomplished without increasing the risk of worker safety since PFKs and HF's are of low toxicity, are non-flammable and are very innocuous materials. Surface films produced with the PFKs and HFKs are more stable to higher temperatures than those formed with sulfur dioxide, enabling work with higher melt temperatures (for additional alloys and more complex casting parts).

Dwg.0/0

TECH WO 200183836 A2UPTX: 20020130

TECHNOLOGY FOCUS - METALLURGY - Preferred Component: The gaseous mixture contains a carrier gas such as air, carbon dioxide, argon, and/or nitrogen.

Preferred Material: The magnesium is in the form of ingots or castings.

TECHNOLOGY FOCUS - ORGANIC CHEMISTRY - Preferred Compounds: The perfluoroketone (PFK) is 1,1,1,2,4,4,5,5,5-nonafluoro-2-trifluoromethyl-pentan-3-one, 1,1,1,2,4,5,5,5-octafluoro-2,4 di(trifluoromethyl)-pentan-3-one, 1,1,1,2,4,4,5,5,6,6,6-undecafluoro-2- trifluoromethyl-hexan-3-one, 1,1,1,2,4,4,5,5,6,6,7,7,7-tridecafluoro-2-trifluoromethyl-heptan-3-one, 1,1,1,3,3,4,4,5,5,6,6,7,7,8,8,8-octafluoro-octan-2-one, 1,1,1,2,2,4,4,5,5,6,6,6-dodecafluoro-hexan-3-one, 1,1,1,3,4,4,4,4-heptafluoro-3-trifluoromethyl-butan-2-one, perfluorocyclohexanone and their mixtures.

The **hydrofluoroketone** (HFK) is  $\text{HCF}_2\text{CF}_2\text{C}(\text{O})\text{CF}(\text{CF}_3)_2$ ,  $\text{CF}_3\text{C}(\text{O})\text{CH}_2\text{C}(\text{O})\text{CF}_3$ ,  $\text{C}_2\text{H}_5\text{C}(\text{O})\text{CF}(\text{CF}_3)_2$ ,  $\text{CF}_2\text{CF}_2\text{C}(\text{O})\text{CH}_3$ ,  $(\text{CF}_3)_2\text{CFC}(\text{O})\text{CH}_3$ ,  $\text{CF}_3\text{CF}_2\text{C}(\text{O})\text{CHF}_2$ ,  $\text{CF}_3\text{CF}_2\text{C}(\text{O})\text{CH}_2\text{F}$ ,  $\text{CF}_3\text{CF}_2\text{C}(\text{O})\text{CH}_2\text{CF}_3$ ,  $\text{CF}_3\text{CF}_2\text{C}(\text{O})\text{CH}_2\text{CH}_3$ ,  $\text{CF}_3\text{CF}_2\text{C}(\text{O})\text{CH}_2\text{CHF}_2$ ,  $\text{CF}_3\text{CF}_2\text{C}(\text{O})\text{CH}_2\text{CHF}_2$ ,  $\text{CF}_3\text{CF}_2\text{C}(\text{O})\text{CH}_2\text{CH}_2\text{F}$ ,  $\text{CF}_3\text{CF}_2\text{C}(\text{O})\text{CHFCH}_3$ ,  $\text{CF}_3\text{CF}_2\text{C}(\text{O})\text{CHFCH}_2\text{F}$ ,  $\text{CF}_3\text{CF}_2\text{C}(\text{O})\text{CF}_2\text{CH}_2\text{F}$ ,  $(\text{CF}_3)_2\text{CFC}(\text{O})\text{CHF}_2$ ,  $(\text{CF}_3)_2\text{CFC}(\text{O})\text{CH}_2\text{F}$ ,  $\text{CF}_3\text{CF}(\text{CH}_2\text{F})\text{C}(\text{O})\text{CHF}_2$ ,  $\text{CH}_3\text{CF}(\text{CH}_2\text{F})\text{C}(\text{O})\text{CH}_2\text{F}$ , and/or  $\text{CF}_3\text{CF}(\text{CH}_2\text{F})\text{C}(\text{O})\text{CF}_3$ , and their mixtures..

ABEX WO 200183836 A2UPTX: 20020130

EXAMPLE - Pure magnesium (3 kg) was heated to 680 degreesC and two cover gases were continuously applied separately to the surface of the molten magnesium. The invented method used 1,1,1,2,4,4,5,5,5-nonafluoro-2-trifluoromethyl-pentan-3-one (perfluorinated ketone) as cover gas, while the comparative process used methoxy nonafluorobutane. The cover gases were evaluated at 1% and at decreasing volumetric concentrations in air. At all concentrations, the perfluorinated ketone produced a film visually thinner and more elastic than that produced by the comparative process. The film produced was stable and did not change appearance over at least 30 minutes. This was in contrast to the test using methoxy nonafluorobutane, where metal burning was noted when the cover gas concentration was reduced to 625 ppm.

FS CPI GMPI

FA AB; DCN  
MC CPI: E10-F02A3; E10-F02C; M25-G16

L49 ANSWER 3 OF 4 WPIX COPYRIGHT 2004 THOMSON DERWENT on STN  
AN 2001-070604 [08] WPIX  
DNN N2001-053480 DNC C2001-019551  
TI Cover gas composition for protecting **molten magnesium**  
(alloy) comprises fluorine containing inhibiting agent and carrier gas.  
DC E16 M22 P35 P53  
IN BAKER, P W; CASHION, S P; FROST, M T; KORN, C J; RICKETTS, N J  
PA (CAST-N) CAST CENT PTY LTD  
CYC 94

*applicants*

PI WO 2000064614 A1 20001102 (200108)\* EN 20 B22D021-04  
RW: AT BE CH CY DE DK EA ES FI FR GB GH GM GR IE IT KE LS LU MC MW NL  
OA PT SD SE SL SZ TZ UG ZW  
W: AE AG AL AM AT AU AZ BA BB BG BR BY CA CH CN CR CU CZ DE DK DM DZ  
EE ES FI GB GD GE GH GM HR HU ID IL IN IS JP KE KG KP KR KZ LC LK  
LR LS LT LU LV MA MD MG MK MN MW MX NO NZ PL PT RO RU SD SE SG SI  
SK SL TJ TM TR TT TZ UA UG US UZ VN YU ZA ZW

AU 2000040930 A 20001110 (200109)  
BR 2000010137 A 20020122 (200216) B22D021-04  
NO 2001005264 A 20011221 (200221) B22D000-00  
EP 1204499 A1 20020515 (200239) EN B22D021-04  
R: AL AT BE CH CY DE DK ES FI FR GB GR IE IT LI LT LU LV MC MK NL PT  
RO SE SI

CZ 2001003817 A3 20020515 (200241) B22D021-04  
KR 2002011397 A 20020208 (200255) B22D021-04  
HU 2002000990 A2 20020729 (200258) B22D021-04  
CN 1352583 A 20020605 (200261) B22D021-04  
SK 2001001556 A3 20020806 (200261) B22D021-04  
JP 2002541999 W 20021210 (200301) 18 A62C003-06  
ZA 2001008862 A 20021224 (200309) 24 B22D000-00  
TW 500805 A 20020901 (200334) C22B026-22  
NZ 515084 A 20031031 (200380) B22D021-04  
AU 766844 B 20031023 (200381) B22D021-04  
MX 2001010941 A1 20030601 (200417) A62D001-02

ADT WO 2000064614 A1 WO 2000-AU393 20000428; AU 2000040930 A AU 2000-40930  
20000428; BR 2000010137 A BR 2000-10137 20000428, WO 2000-AU393 20000428;  
NO 2001005264 A WO 2000-AU393 20000428, NO 2001-5264 20011026; EP 1204499  
A1 EP 2000-920274 20000428, WO 2000-AU393 20000428; CZ 2001003817 A3 WO  
2000-AU393 20000428, CZ 2001-3817 20000428; KR 2002011397 A KR 2001-713855  
20011029; HU 2002000990 A2 WO 2000-AU393 20000428, HU 2002-990 20000428;  
CN 1352583 A CN 2000-808146 20000428; SK 2001001556 A3 WO 2000-AU393  
20000428, SK 2001-1556 20000428; JP 2002541999 W JP 2000-613596 20000428,  
WO 2000-AU393 20000428; ZA 2001008862 A ZA 2001-8862 20011026; TW 500805 A  
TW 2000-108130 20000504; NZ 515084 A NZ 2000-515084 20000428, WO  
2000-AU393 20000428; AU 766844 B AU 2000-40930 20000428; MX 2001010941 A1  
WO 2000-AU393 20000428, MX 2001-10941 20011026

FDT AU 2000040930 A Based on WO 2000064614; BR 2000010137 A Based on WO  
2000064614; EP 1204499 A1 Based on WO 2000064614; CZ 2001003817 A3 Based  
on WO 2000064614; HU 2002000990 A2 Based on WO 2000064614; SK 2001001556  
A3 Based on WO 2000064614; JP 2002541999 W Based on WO 2000064614; NZ  
515084 A Based on WO 2000064614; AU 766844 B Previous Publ. AU 2000040930,  
Based on WO 2000064614; MX 2001010941 A1 Based on WO 2000064614

PRAI AU 1999-15 19990428

IC ICM A62C003-06; A62D001-02; B22D000-00

ICS B22D021-02; C22B004-02; C22C001-02; C22C023-00

ICA B22D021-04; C22B026-22

AB WO 2000064614 A UPAB: 20010207

NOVELTY - Cover gas composition comprises a fluorine containing inhibiting agent and a carrier gas. Each component has a Global Warming Potential (GWP) of less than 5000.

DETAILED DESCRIPTION - INDEPENDENT CLAIMS are also included for:

- (a) a method of protecting **molten magnesium** (alloy);
- (b) use of the inhibiting agent for preventing or minimizing oxidation of **molten magnesium** (alloy);
- (c) a method of extinguishing a magnesium (alloy) fire.

USE - For use as cover gases for protecting **molten magnesium** (alloys).

ADVANTAGE - The inventive cover gas composition provides good industrial scale protection of **molten magnesium** at a lower concentration.

Dwg.0/0

TECH WO 200064614 A1UPTX: 20010207

TECHNOLOGY FOCUS - INORGANIC CHEMISTRY - Preferred Gas: The carrier gas is carbon dioxide, argon, nitrogen, and/or preferably dry air.

TECHNOLOGY FOCUS - ORGANIC CHEMISTRY - Preferred Materials: The inhibiting agent is **difluoromethane, pentafluoromethane, difluoroethane, heptafluoropropane, methoxy-nonafluorobutane, ethoxy-nonafluorobutane, dihydrodecafluoropentane**, and/or preferably 1,1,1,2-tetrafluoroethane.

Preferred Properties: The inhibiting agent has a boiling point of less than 100 degrees C and has no ozone depletion potential.

The composition has a GWP of less than 3000 (preferably less than 1500).

Preferred Composition: The composition contains inhibiting agent (less than 1, preferably less than 0.1 vol.%).

ABEX WO 200064614 A1UPTX: 20010207

EXAMPLE - A crucible furnace containing 100g of molten pure magnesium at 680 degrees C was blanketed with a gaseous composition consisting of 1,1,1,2-tetrafluoroethane (0.02 volume%) and dry air (balance). Good molten protection was observed, with the formation of a thin protective surface film.

FS CPI GMPI

FA AB; DCN

MC CPI: E10-H04A3; E31-H03; E31-J; E31-N05C; M22-G

L49 ANSWER 4 OF 4 WPIX COPYRIGHT 2004 THOMSON DERWENT on STN

AN 1996-171062 [17] WPIX

DNC C1996-053941

TI Catalyst preparation for production of 1,1,1,2-tetra fluoroethane - by producing

admixture of compsn. comprising chromium hydroxide hydrate and e.g. magnesium chloride with aqueous metal salt, reacting in liquid phase, etc..

DC E16 J04

IN KIM, H; LEE, B; PARK, K; KIM, H S; LEE, B G; PARK, K Y

PA (KOAD) KOREA INST SCI & TECHNOLOGY; (KOAD) KOREA ADV INST SCI & TECHNOLOGY

CYC 2

PI US 5500400 A 19960319 (199617)\* 4 B01J023-26

KR 9616683 B1 19961220 (199931) B01J023-26

ADT US 5500400 A US 1994-302636 19940908; KR 9616683 B1 KR 1993-27091 19931209

PRAI KR 1993-27091 19931209

IC ICM B01J023-26

ICS B01J023-78; C07C019-08

AB US 5500400 A UPAB: 19960428

The catalyst for the production of 1,1,1,2-tetrafluoroethane (HFC-134a) is prepared by: (a) preparing an admixture of a compsn.

comprising Cr hydroxide hydrate and MgCl<sub>2</sub> or CaCl<sub>2</sub> in a weight ratio of Cr to Mg or Ca of 1:0.3-10 with an aqueous metal salt solution from Ce, Mg, Ni and Al chlorides, (b) reacting the mixture in the liquid phase with an aqueous HF solution

to give a paste; and (c) sintering the paste.

Also claimed is a catalyst for the production of HFC-134a represented by formula Cr<sub>a</sub>L<sub>b</sub>M<sub>c</sub>O<sub>x</sub>F<sub>y</sub> (I), where: L = Mg or Ca; M = one of Ce, Ni, Zn, and Al; and a, b, c, x and y = 0-2.

USE - Used in the production of HFC-134a for use in refrigerators, car cooling systems, etc., by reaction of HCFC-133a with HF.

ADVANTAGE - The catalyst has improved selectivity and durability, and remains active over extended periods. Pre-treatment with HF solution is not required, and the supply of O<sub>2</sub> to reduce catalyst deactivation is also not required, thus reducing by-prod. formation, HCl oxidation, and water formation, lowering corrosion in the appts..

Dwg.0/0

FS CPI

FA AB; DCN

MC CPI: E10-H03A3; E35-C; E35-P; E35-W; J04-E04; N01-B; N01-C; N03-A; N03-D01; N03-F; N06-E

=> FILE HCAPLU

FILE 'HCAPLUS' ENTERED AT 11:56:28 ON 20 SEP 2004

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FILE COVERS 1907 - 20 Sep 2004 VOL 141 ISS 13

FILE LAST UPDATED: 19 Sep 2004 (20040919/ED)

This file contains CAS Registry Numbers for easy and accurate substance identification.

=> D QUE L37

L22 1 SEA FILE=REGISTRY ABB=ON 12634-54-7  
 L23 13 SEA FILE=REGISTRY ABB=ON (124-38-9/BI OR 12634-54-7/BI OR 138495-42-8/BI OR 163702-05-4/BI OR 219484-64-7/BI OR 33660-75-2/BI OR 354-33-6/BI OR 7439-95-4/BI OR 7440-37-1/BI OR 75-10-5/BI OR 75-37-6/BI OR 7727-37-9/BI OR 811-97-2/BI)  
 L24 8 SEA FILE=REGISTRY ABB=ON L23 AND 1-20/F  
 L25 1 SEA FILE=REGISTRY ABB=ON MAGNESIUM/CN  
 L28 1 SEA FILE=REGISTRY ABB=ON "1,1,1,2-TETRAFLUOROETHANE"/CN  
 L32 1 SEA FILE=REGISTRY ABB=ON 163702-08-7  
 L34 1543680 SEA FILE=HCAPLUS ABB=ON L22 OR L25 OR (MG OR MAGNESIUM)  
 L35 19560 SEA FILE=HCAPLUS ABB=ON L34 (6A) (MELT? OR MOLTEN? OR LIQ?)  
 L36 8793 SEA FILE=HCAPLUS ABB=ON L24 OR L28 OR L32

L37 8 SEA FILE=HCAPLUS ABB=ON L35.AND L36

=> FILE COMPENDEX

FILE 'COMPENDEX' ENTERED AT 11:56:41 ON 20 SEP 2004

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FILE LAST UPDATED: 20 SEP 2004 <20040920/UP>

FILE COVERS 1970 TO DATE.

<<< SIMULTANEOUS LEFT AND RIGHT TRUNCATION AVAILABLE IN  
THE BASIC INDEX >>>

=> D QUE L47

L22 1 SEA FILE=REGISTRY ABB=ON 12634-54-7  
L25 1 SEA FILE=REGISTRY ABB=ON MAGNESIUM/CN  
L34 1543680 SEA FILE=HCAPLUS ABB=ON L22 OR L25 OR (MG OR MAGNESIUM)  
L39 6125 SEA FILE=WPIX ABB=ON L34 (6A) (MELT? OR MOLTEN? OR LIQ?)  
L42 3 SEA FILE=WPIX ABB=ON L39 AND (DIFLUOROMETHANE OR PENTAFLUOROET  
HANE OR DIFLUOROETHANE OR HEPTAFLUOROPROPANE OR (METHOXY OR  
ETHOXY) (W)NONAFLUOROBUTANE OR DIHYDRODECAFLUOROPENTANE OR  
TETRAFLUOROETHANE)  
L43 0 SEA FILE=WPIX ABB=ON L39 AND (HYDROFLUOROETHER# OR HYDRO(W) FLU  
ORO(W)ETHER# OR HYDROFLUORO(W)ETHER#)  
L45 1 SEA FILE=COMPENDEX ABB=ON L42 OR L43  
L46 1 SEA FILE=COMPENDEX ABB=ON L39 AND HYDROFLUORO?  
L47 1 SEA FILE=COMPENDEX ABB=ON L45 OR L46

=> FILE METADEX

FILE 'METADEX' ENTERED AT 11:57:02 ON 20 SEP 2004

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FILE LAST UPDATED: 30 AUG 2004 <20040830/UP>

FILE COVERS 1966 TO DATE.

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=> D QUE L50

L22 1 SEA FILE=REGISTRY ABB=ON 12634-54-7  
L25 1 SEA FILE=REGISTRY ABB=ON MAGNESIUM/CN  
L34 1543680 SEA FILE=HCAPLUS ABB=ON L22 OR L25 OR (MG OR MAGNESIUM)  
L39 6125 SEA FILE=WPIX ABB=ON L34 (6A) (MELT? OR MOLTEN? OR LIQ?)  
L42 3 SEA FILE=WPIX ABB=ON L39 AND (DIFLUOROMETHANE OR PENTAFLUOROET  
HANE OR DIFLUOROETHANE OR HEPTAFLUOROPROPANE OR (METHOXY OR  
ETHOXY) (W)NONAFLUOROBUTANE OR DIHYDRODECAFLUOROPENTANE OR  
TETRAFLUOROETHANE)  
L43 0 SEA FILE=WPIX ABB=ON L39 AND (HYDROFLUOROETHER# OR HYDRO(W) FLU  
ORO(W)ETHER# OR HYDROFLUORO(W)ETHER#)  
L44 3 SEA FILE=WPIX ABB=ON L42 OR L43  
L46 1 SEA FILE=COMPENDEX ABB=ON L39 AND HYDROFLUORO?  
L50 3 SEA FILE=METADEX ABB=ON L44 OR L46

=> FILE NTIS

FILE 'NTIS' ENTERED AT 11:57:28 ON 20 SEP 2004

KATHLEEN FULLER EIC 1700 REMSEN 4B28 571/272-2505



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FILE LAST UPDATED: 18 SEP 2004 <20040918/UP>  
FILE COVERS 1964 TO DATE.

<<<SIMULTANEOUS LEFT AND RIGHT TRUNCATION AVAILABLE IN  
THE BASIC INDEX (/BI) >>>

=> D QUE L51

L22 1 SEA FILE=REGISTRY ABB=ON 12634-54-7  
L25 1 SEA FILE=REGISTRY ABB=ON MAGNESIUM/CN  
L34 1543680 SEA FILE=HCAPLUS ABB=ON L22 OR L25 OR (MG OR MAGNESIUM)  
L39 6125 SEA FILE=WPIX ABB=ON L34 (6A) (MELT? OR MOLTEN? OR LIQ?)  
L42 3 SEA FILE=WPIX ABB=ON L39 AND (DIFLUOROMETHANE OR PENTAFLUOROET  
HANE OR DIFLUOROETHANE OR HEPTAFLUOROPROPANE OR (METHOXY OR  
ETHOXY) (W)NONAFLUOROBUTANE OR DIHYDRODECAFLUOROPENTANE OR  
TETRAFLUOROETHANE)  
L43 0 SEA FILE=WPIX ABB=ON L39 AND (HYDROFLUOROETHER# OR HYDRO(W)FLU  
ORO(W)ETHER# OR HYDROFLUORO(W)ETHER#)  
L44 3 SEA FILE=WPIX ABB=ON L42 OR L43  
L46 1 SEA FILE=COMPENDEX ABB=ON L39 AND HYDROFLUORO?  
L51 0 SEA FILE=NTIS ABB=ON L44 OR L46

=> FILE JICST

FILE 'JICST-EPLUS' ENTERED AT 11:57:42 ON 20 SEP 2004  
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FILE COVERS 1985 TO 13 SEP 2004 (20040913/ED)

THE JICST-EPLUS FILE HAS BEEN RELOADED TO REFLECT THE 1999 CONTROLLED  
TERM (/CT) THESAURUS RELOAD.

=> D QUE L52

L22 1 SEA FILE=REGISTRY ABB=ON 12634-54-7  
L25 1 SEA FILE=REGISTRY ABB=ON MAGNESIUM/CN  
L34 1543680 SEA FILE=HCAPLUS ABB=ON L22 OR L25 OR (MG OR MAGNESIUM)  
L39 6125 SEA FILE=WPIX ABB=ON L34 (6A) (MELT? OR MOLTEN? OR LIQ?)  
L42 3 SEA FILE=WPIX ABB=ON L39 AND (DIFLUOROMETHANE OR PENTAFLUOROET  
HANE OR DIFLUOROETHANE OR HEPTAFLUOROPROPANE OR (METHOXY OR  
ETHOXY) (W)NONAFLUOROBUTANE OR DIHYDRODECAFLUOROPENTANE OR  
TETRAFLUOROETHANE)  
L43 0 SEA FILE=WPIX ABB=ON L39 AND (HYDROFLUOROETHER# OR HYDRO(W)FLU  
ORO(W)ETHER# OR HYDROFLUORO(W)ETHER#)  
L44 3 SEA FILE=WPIX ABB=ON L42 OR L43  
L46 1 SEA FILE=COMPENDEX ABB=ON L39 AND HYDROFLUORO?  
L52 0 SEA FILE=JICST-EPLUS ABB=ON L44 OR L46

=> FILE JAPIO

FILE 'JAPIO' ENTERED AT 11:57:55 ON 20 SEP 2004  
COPYRIGHT (C) 2004 Japanese Patent Office (JPO)- JAPIO

FILE LAST UPDATED: 3 SEP 2004 <20040903/UP>  
FILE COVERS APR 1973 TO APRIL 30, 2004

<<< GRAPHIC IMAGES AVAILABLE >>>

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=> D QUE L53

L22 1 SEA FILE=REGISTRY ABB=ON 12634-54-7  
 L25 1 SEA FILE=REGISTRY ABB=ON MAGNESIUM/CN  
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 L39 6125 SEA FILE=WPIX ABB=ON L34 (6A) (MELT? OR MOLTEN? OR LIQ?)  
 L42 3 SEA FILE=WPIX ABB=ON L39 AND (DIFLUOROMETHANE OR PENTAFLUOROETHANE OR DIFLUOROETHANE OR HEPTAFLUOROPROPANE OR (METHOXY OR ETHOXY) (W)NONAFLUOROBUTANE OR DIHYDRODECAFLUOROPENTANE OR TETRAFLUOROETHANE)  
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 L46 1 SEA FILE=COMPENDEX ABB=ON L39 AND HYDROFLUORO?  
 L53 0 SEA FILE=JAPIO ABB=ON L44 OR L46

=> DUP REM L37 L47 L50

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 COPYRIGHT (c) 2004 Cambridge Scientific Abstracts (CSA)  
 PROCESSING COMPLETED FOR L37  
 PROCESSING COMPLETED FOR L47  
 PROCESSING COMPLETED FOR L50  
L54 9 DUP REM L37 L47 L50 (3 DUPLICATES REMOVED)

=> D L54 ALL HITSTR 1-9

L54 ANSWER 1 OF 9 HCAPLUS COPYRIGHT 2004 ACS on STN  
 AN 2003:409780 HCAPLUS  
 DN 139:10048  
 ED Entered STN: 29 May 2003  
 TI **Magnesium melt protection at Magnesium**  
 Elektron using HFC-134a  
 AU Lyon, Paul; Rogers, Philip D.; King, John F.; Cashion, Simon P.; Ricketts, Nigel J.  
 CS Development Dept., Magnesium Elektron, Manchester, M27 8DD, UK  
 SO Magnesium Technology 2003, Proceedings of the Symposium held during the 2003 TMS Annual Meeting, San Diego, CA, United States, ar. 2-6, 2003 (2003), 11-14. Editor(s): Kaplan, Howard I. Publisher: Minerals, Metals & Materials Society, Warrendale, Pa.  
 CODEN: 69DYJV; ISBN: 0-87339-533-6  
 DT Conference  
 LA English  
 CC 56-2 (Nonferrous Metals and Alloys)  
 AB Several workers are actively seeking alternatives to SF6. CAST/AMC have patented the use of a HFC gas HFC-134a. This gas has a GWP 95% lower than SF6. Magnesium Elektron (MEL) and CAST have collaborated on the use of HFC-134a to achieve successful production plant trials for ingot manufacture  
 This

paper provides details and results of those plant trials at Magnesium Elektron. Magnesium Elektron ingot production is currently based on either SF6 or SO2 as the active gas. Comparative evaluation of HFC-134a included, development of gas mixing equipment, distribution optimization, optimization of concentration and flow rates and assessment of breakdown products

and potential implications. The results, from over 150 production-scale Mg-Al-Zn melts, demonstrated that HFC-134a could offer equal protection to SF6 or SO2. Use of HFC-134a by Magnesium Elektron for com. production is proposed.

ST magnesium casting melt protection fluorinated hydrocarbon

IT Melting

(alloy; magnesium melt protection in casting with HFC-134a)

IT Hydrocarbons, uses

RL: NUU (Other use, unclassified); USES (Uses)

(fluoro; magnesium melt protection in casting with HFC-134a)

IT Casting of metals

Controlled atmospheres

(magnesium melt protection in casting with HFC-134a)

IT 811-97-2, HFC-134a

RL: NUU (Other use, unclassified); USES (Uses)

(magnesium melt protection in casting with HFC-134a)

IT 135860-09-2, AZ91D

RL: PEP (Physical, engineering or chemical process); PYP (Physical process); PROC (Process)

(magnesium melt protection in casting with HFC-134a)

RE.CNT 9 THERE ARE 9 CITED REFERENCES AVAILABLE FOR THIS RECORD  
RE

(1) Bartos, S; Magnesium Technology 2000 2000, P83

(2) Cashion, S; Paper presented at the 8th Annual IMA Magnesium in Automotive Seminar 2000

(3) Cashion, S; PhD Thesis, University of Queensland 1999

(4) Erickson, S; Foundry Management and Technology 1998, V126(6), P38

(5) Fruehling, J; PhD thesis, University of Michigan 1970

(6) Gjestland, H; The 3rd International Magnesium Conference 1996, P33

(7) Granier, C; Scientific Assessment of Ozone Depletion 1998

(8) Reimers, H; US 1972317 1934 HCAPLUS

(9) Ricketts, N; Magnesium Technology 2001 2001, P31 HCAPLUS

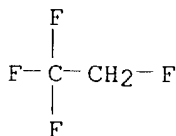
IT 811-97-2, HFC-134a

RL: NUU (Other use, unclassified); USES (Uses)

(magnesium melt protection in casting with HFC-134a)

RN 811-97-2 HCAPLUS

CN Ethane, 1,1,1,2-tetrafluoro- (6CI, 7CI, 8CI, 9CI) (CA INDEX NAME)



L54 ANSWER 2 OF 9 HCAPLUS COPYRIGHT 2004 ACS on STN  
 AN 2002:403070 HCAPLUS  
 DN 137:220130  
 ED Entered STN: 30 May 2002  
 TI Characterisation of the surface films formed on **molten magnesium** in different protective atmospheres  
 AU Pettersen, Gunnar; Ovrelid, E.; Tranell, G.; Fenstad, J.; Gjestland, H.  
 CS SINTEF Materials Technology, Trondheim, N-7465, Norway  
 SO Materials Science & Engineering, A: Structural Materials: Properties, Microstructure and Processing (2002), A332(1-2), 285-294  
 CODEN: MSAPE3; ISSN: 0921-5093  
 PB Elsevier Science B.V.  
 DT Journal  
 LA English  
 CC 56-2 (Nonferrous Metals and Alloys)  
 AB **Molten Mg** oxidizes rapidly during casting and handling unless it is protected by an atmospheric that stabilizes the surface. In this article, results from the anal. of **Mg melt** surfaces exposed to SO<sub>2</sub> and different fluorine-containing atmospheres are reported. The microstructure of the surface films, formed during controlled exposure in laboratory scale expts., have been characterized using X-ray diffraction (XRD), electron probe microanal. (EPMA) and transmission electron microscopy (TEM). Both SO<sub>2</sub> and the fluorine-containing gases were found to protect the melt from burning and vaporization in oxidizing atmospheres. The protected surfaces generally had a shiny metallic appearance, but turned dull gray after extended exposure to high concns. of fluorine containing gases. All the surface films initially consisted of small crystallites of MgO forming a thin continuous film. This film was found to contain some sulfur when the melt was protected by SO<sub>2</sub>, while fluorine was the only element detected in the oxide when SF<sub>6</sub> or other fluorine-containing gases were used for protection. With increasing exposure time, the films gradually grew thicker and the fluorine/oxygen-ratio of the films formed in fluorine-containing atmospheres increased. Finally, after long term exposure to fluorine containing atmospheres, the thermodynamically stable MgF<sub>2</sub>-phase was formed. In a N<sub>2</sub> atmosphere, SO<sub>2</sub> and SF<sub>6</sub>-addns. did not protect the magnesium, indicating that a rapid initial formation of MgO is necessary to obtain protective films.  
 ST controlled atm film formation **molten magnesium melt** oxidn prevention  
 IT Casting of metals  
 IT Controlled atmospheres  
 (characterization of the surface films formed on **molten magnesium** in different protective atmospheres)  
 IT 7439-95-4P, Magnesium, preparation  
 RL: IMF (Industrial manufacture); PRP (Properties); RCT (Reactant); TEM (Technical or engineered material use); PREP (Preparation); RACT (Reactant or reagent); USES (Uses)  
 (characterization of the surface films formed on **molten magnesium** in different protective atmospheres)  
 IT 7446-09-5, Sulfur dioxide, uses  
 RL: TEM (Technical or engineered material use); USES (Uses)  
 (characterization of the surface films formed on **molten magnesium** in different protective atmospheres)  
 IT 75-28-5, Isobutane 76-19-7 **811-97-2** 454695-19-3, ICEON 49  
 RL: NUU (Other use, unclassified); USES (Uses)  
 (gas atmosphere containing; characterization of the surface films formed on **molten magnesium** in different protective atmospheres)  
 IT 2551-62-4, Sulfur hexafluoride 7783-54-2, Trifluoroamine

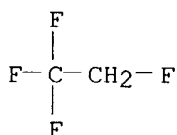
RL: NUU (Other use, unclassified); USES (Uses)  
 (gas atmosphere; characterization of the surface films formed on  
**molten magnesium** in different protective atmospheres)

IT 1309-48-4, Magnesium oxide (MgO), properties 7783-40-6, Magnesium  
 fluoride (MgF<sub>2</sub>) 12032-36-9, Magnesium sulfide (MgS)  
 RL: FMU (Formation, unclassified); PRP (Properties); TEM (Technical or  
 engineered material use); FORM (Formation, nonpreparative); USES (Uses)  
 (protective phase; characterization of the surface films formed on  
**molten magnesium** in different protective atmospheres)

RE.CNT 11 THERE ARE 11 CITED REFERENCES AVAILABLE FOR THIS RECORD  
 RE

- (1) Cashion, S; PhD thesis, Department of Mining, Minerals and Materials  
 Engineering, University of Queensland 1998
- (2) Couling, S; Light Metal Age 1977, P12 HCAPLUS
- (3) Couling, S; Proc of the 36th Annual World Conference on Magnesium 1979, P54  
 HCAPLUS
- (4) Fruehling, J; PhD thesis, The University of Michigan 1970
- (5) Gjestland, H; Proc of the Third Int Magnesium Conference 1996, P33
- (6) Hinz, W; The American Mineralogist 1960, V45, P1198 HCAPLUS
- (7) JCPDS-International Centre for Diffraction Data; Powder Diffraction File,  
 Sets 1-44
- (8) Lindseth, I; Proc 2nd Int Symp on Aluminium Surface Science and Technology  
 (ASST2000) 2000, P553 HCAPLUS
- (9) Putz, H; Computational Materials Science 1998, V11, P309 HCAPLUS
- (10) Reimers, H; US 1972317 1934 HCAPLUS
- (11) Roine, A; HSC Chemical Reaction and Equilibrium Software with extensive  
 Thermochemical Database, version 2.03 1994

IT **811-97-2**  
 RL: NUU (Other use, unclassified); USES (Uses)  
 (gas atmosphere containing; characterization of the surface films formed on  
**molten magnesium** in different protective atmospheres)  
 RN 811-97-2 HCAPLUS  
 CN Ethane, 1,1,1,2-tetrafluoro- (6CI, 7CI, 8CI, 9CI) (CA INDEX NAME)



L54 ANSWER 3 OF 9 HCAPLUS COPYRIGHT 2004 ACS on STN DUPLICATE 1  
 AN 2002:438031 HCAPLUS  
 DN 137:173739  
 ED Entered STN: 11 Jun 2002  
 TI Replacing SF<sub>6</sub> with the hydrofluorocarbon gas HFC-134a for  
**magnesium melt** protection  
 AU Cashion, Simon P.; Ricketts, Nigel J.  
 CS Cooperative Research Centre for Cast Metals Manufacturing (CAST)  
 Department of Mining, Minerals and Materials Engineering, The University  
 of Queensland, St. Lucia, 4072, Australia  
 SO Greenhouse Gases in the Metallurgical Industries: Policies, Abatement and  
 Treatment, Proceedings of the International Symposium on Greenhouse Gases  
 in the Metallurgical Industries: Policies, Abatement and Treatment,  
 Toronto, ON, Canada, Aug. 26-29, 2001 (2001), 315-324. Editor(s):  
 Pickles, Chris A. Publisher: Canadian Institute of Mining, Metallurgy and  
 Petroleum, Montreal, Que.

CODEN: 69CROA; ISBN: 1-894475-15-1

DT Conference

LA English

CC 59-2 (Air Pollution and Industrial Hygiene)

Section cross-reference(s): 56

AB Since the early 1970s, sulfur hexafluoride (SF6) has been used to protect **molten magnesium** from oxidation. However the use of SF6 is coming under increased scrutiny. It is now recognized as a very potent greenhouse gas, and has been targeted for phase-out under the Kyoto protocol. In the search for an alternative to SF6, the Cooperative Research Center for Cast Metals Manufacturing (CAST) in Australia is developing a cover gas system based upon the hydrofluorocarbon gas 1,1,1,2-tetrafluoroethane (HFC-134a). This gas has a global warming potential 17 times lower than SF6 and has been shown to be an effective replacement for SF6 in many applications. Recent experience with CAST's new cover gas system in laboratory and plant trials is discussed as well as the global implications of replacing SF6 with HFC-134a.

ST air pollution greenhouse gas sulfur hexafluoride tetrafluoroethane magnesium metallurgy

IT Air pollution

Greenhouse gases

(replacing the greenhouse gas SF6 with tetrafluoroethane for **magnesium melt** protection)

IT 135860-09-2P, Az91d

RL: IMF (Industrial manufacture); PREP (Preparation)

(AZ91D, protection of molten; replacing the greenhouse gas SF6 with tetrafluoroethane for **magnesium melt** protection)

IT 7664-39-3P, Hydrogen fluoride, biological studies

RL: ADV (Adverse effect, including toxicity); BYP (Byproduct); BIOL (Biological study); PREP (Preparation)

(decomposition product of HFC-134a; replacing the greenhouse gas SF6 with tetrafluoroethane for **magnesium melt** protection)

IT 7439-95-4P, Magnesium, preparation

RL: IMF (Industrial manufacture); PREP (Preparation)

(protection of **molten**; replacing the greenhouse gas SF6 with tetrafluoroethane for **magnesium melt** protection)

IT 811-97-2, Ethane, 1,1,1,2-tetrafluoro- 2551-62-4, Sulfur fluoride (SF6), (OC-6-11)-

RL: NUU (Other use, unclassified); POL (Pollutant); OCCU (Occurrence); USES (Uses)

(replacing the greenhouse gas SF6 with tetrafluoroethane for **magnesium melt** protection)

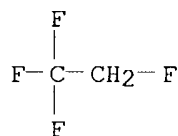
RE.CNT 13 THERE ARE 13 CITED REFERENCES AVAILABLE FOR THIS RECORD  
RE

- (1) Australian Bureau Of Statistics; Motor Vehicle Census 1998, 9309.0, P3
- (2) Bartos, S; Magnesium Technology 2000 2000, P83
- (3) Cashion, S; Paper presented at the 8th Annual IMA Magnesium in Automotive Seminar 2000
- (4) Cashion, S; Ph D Thesis, University of Queensland 1999
- (5) Erickson, S; Foundry Management and Technology 1998, V126(6), P38
- (6) Fruehling, J; PhD thesis, University of Michigan 1970
- (7) Gjestland, H; The 3rd International Magnesium Conference 1996, P33
- (8) Granier, C; Scientific Assessment of Ozone Depletion:1998
- (9) Mellerud, T; The 56th IMA Annual Meeting 1999
- (10) Reimers, H; US 1972317 1934 HCAPLUS
- (11) Ricketts, N; WO PCTAU0000393 2000
- (12) Ricketts, N; Magnesium Technology 2001 2001, P31 HCAPLUS
- (13) Wilkenfeld, G; Analysis of Trends and Greenhouse Indicators 1990 to 1998 2000, P41

IT 7439-95-4P, Magnesium, preparation  
 RL: IMF (Industrial manufacture); PREP (Preparation)  
 (protection of **molten**; replacing the greenhouse gas SF6 with  
 tetrafluoroethane for **magnesium melt** protection)  
 RN 7439-95-4 HCAPLUS  
 CN Magnesium (8CI, 9CI) (CA INDEX NAME)

Mg

IT 811-97-2, Ethane, 1,1,1,2-tetrafluoro-  
 RL: NUU (Other use, unclassified); POL (Pollutant); OCCU (Occurrence);  
 USES (Uses)  
 (replacing the greenhouse gas SF6 with tetrafluoroethane for  
**magnesium melt** protection)  
 RN 811-97-2 HCAPLUS  
 CN Ethane, 1,1,1,2-tetrafluoro- (6CI, 7CI, 8CI, 9CI) (CA INDEX NAME)



L54 ANSWER 4 OF 9 HCAPLUS COPYRIGHT 2004 ACS on STN DUPLICATE 2  
 AN 2001:272575 HCAPLUS  
 DN 134:343042  
 ED Entered STN: 18 Apr 2001  
 TI Hydrofluorocarbons as a replacement for sulphur hexafluoride in magnesium  
 processing  
 AU Ricketts, Nigel J.; Cashion, Simon P.  
 CS Cooperative Research Centre for Cast Metals Manufacturing (CAST),  
 Pullenvale, 4069, Australia  
 SO Magnesium Technology 2001, Proceedings of the Symposium held during the  
 TMS Annual Meeting, New Orleans, LA, United States, Feb. 11-15, 2001  
 (2001), 31-36. Editor(s): Hryn, John N. Publisher: Minerals, Metals &  
 Materials Society, Warrendale, Pa.  
 CODEN: 69BDNU  
 DT Conference  
 LA English  
 CC 56-2 (Nonferrous Metals and Alloys)  
 AB One of the major issues for users and potential users of **magnesium**  
 is **magnesium melt** protection. This is particularly so  
 with the use of sulfur hexafluoride coming under increased scrutiny, as it  
 is now recognized as a very potent greenhouse gas. It has been  
 demonstrated that sulfur hexafluoride prevents **molten**  
**magnesium** from oxidation by adding fluorine into the unprotective  
 magnesium oxide surface film. The hydrofluorocarbon gas  
 1,1,1,2-tetrafluoroethane (HFC-134a) has been recently developed for  
**magnesium melt** protection by the Cooperative Research  
 Center for Cast Metals Manufacturing (CAST) in Australia. This gas has a  
 global  
 warming potential 18 times lower than sulfur hexafluoride and has been  
 shown to be an effective replacement for SF6 in many applications.  
 HFC-134a has also been shown to have some unique properties resulting in

more possible applications for **magnesium melt** protection than sulfur hexafluoride. Recent experience with HFC-134a in laboratory and plant trials is discussed as well as the global implications of replacing sulfur hexafluoride with hydrofluorocarbons.

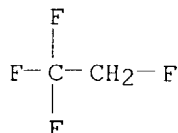
- ST **magnesium alloy melting** hydrofluorocarbon protection  
 IT Casting of metals  
     (hydrofluorocarbons as replacement for sulfur hexafluoride for protection of **molten magnesium** alloys)  
 IT Cast alloys  
     RL: PEP (Physical, engineering or chemical process); PRP (Properties); PROC (Process)  
     (magnesium alloys; hydrofluorocarbons as replacement for sulfur hexafluoride for protection of **molten magnesium** alloys)  
 IT 7439-95-4P, Magnesium, preparation  
     RL: IMF (Industrial manufacture); PREP (Preparation)  
     (hydrofluorocarbons as replacement for sulfur hexafluoride for protection of **molten magnesium** alloys)  
 IT **811-97-2**, 1,1,1,2-Tetrafluoroethane  
     RL: NUU (Other use, unclassified); PEP (Physical, engineering or chemical process); PROC (Process); USES (Uses)  
     (hydrofluorocarbons as replacement for sulfur hexafluoride for protection of **molten magnesium** alloys)  
 IT 135860-09-2, Az91d  
     RL: PEP (Physical, engineering or chemical process); PRP (Properties); PROC (Process)  
     (hydrofluorocarbons as replacement for sulfur hexafluoride for protection of **molten magnesium** alloys)  
 RE.CNT 20 THERE ARE 20 CITED REFERENCES AVAILABLE FOR THIS RECORD  
 RE  
 (1) Anon; Private communications with Solvay und Fluor GmbH  
 (2) Bartos, S; Magnesium Technology 2000 2000, P83  
 (3) Cashion, S; PhD thesis, The University of Queensland 1998  
 (4) Cashion, S; paper presented at the 8th Annual IMA Magnesium in Automotive Seminar 2000  
 (5) Couling, S; 36th Annual World Conference on Magnesium 1979  
 (6) Deinzer, G; Magnesium Alloys and their Applications 1998, P119 HCAPLUS  
 (7) Erickson, S; Foundry Management & Technology 1998, V126(6), P38  
 (8) Erickson, S; International Magnesium Association Technical Committee Report  
 (9) Fruehling, J; PhD thesis, University of Michigan 1970  
 (10) Gjestland, H; Proceedings of the third International Magnesium Conference 1996, P33  
 (11) Harvey, R; Estimates of US emissions of high-global warming potential gases and the costs of reductions, Draft Report 2000  
 (12) Kahn, D; papers presented at Magnesium Alloys and their Applications 1998, P471 HCAPLUS  
 (13) Li, Y; REWARS'99  
 (14) Mellerud, T; Presented at 56th IMA Annual Meeting 1999  
 (15) Peschaux, M; PCT Patent Application, WO 99/43458 1999  
 (16) Pilling, N; Journal of the Institute of Metals 1923, V29, P529  
 (17) Reimers, H; US 1972317 1934 HCAPLUS  
 (18) Revankar, V; 57th Annual World Magnesium Conference 2000, P51  
 (19) Schwarz, W; Emissions and reduction potentials of hydrofluorocarbons, perfluorocarbons and sulphur hexafluoride in Germany 1999  
 (20) Sturges, W; Science V289, P611 HCAPLUS  
 IT **811-97-2**, 1,1,1,2-Tetrafluoroethane  
     RL: NUU (Other use, unclassified); PEP (Physical, engineering or chemical process); PROC (Process); USES (Uses)  
     (hydrofluorocarbons as replacement for sulfur hexafluoride for



protection of molten magnesium alloys)

RN 811-97-2 HCAPLUS

CN Ethane, 1,1,1,2-tetrafluoro- (6CI, 7CI, 8CI, 9CI) (CA INDEX NAME)



L54 ANSWER 5 OF 9 METADEX COPYRIGHT 2004 CSA on STN

AN 2001(12):51-2066 METADEX

TI **Hydrofluorocarbons** as a replacement for SF6.

AU Ricketts, N.J. (CSIRO); Cashion, S.P. (University of Queensland)

SO Magnesium Industry (Sept. 2001) 2, (5), 26-31, Graphs  
ISSN: 1590-959X

DT Journal

CY Italy

LA English

AB One of the major issues for users and potential users of **magnesium** is **magnesium melt** protection. This is particularly so with the use of sulphur hexafluoride coming under increased scrutiny, as it is now recognized as a very potent greenhouse gas. It has been demonstrated that sulphur hexafluoride prevents **molten magnesium** from oxidation by adding fluorine into the unprotective magnesium oxide surface film. The **hydrofluorocarbon** gas 1,1,1,2-tetrafluoroethane (HFC-134a) has been recently developed for **magnesium melt** protection by the Cooperative Research Centre for Cast Metals Manufacturing (CAST) in Australia. This gas has a global warming potential 18 times lower than sulphur hexafluoride and has been shown to be an effective replacement for SF6 in many applications. HFC-134a has also been shown to have some unique properties resulting in more possible applications for **magnesium melt** protection than sulphur hexafluoride.

CC 51 Foundry

CT Journal Article; **Magnesium: Melting**; Oxidation resistance: Processing effects; Combustion: Processing effects; Fluorides: Materials selection; Halocarbons; Sulfur compounds: Materials substitution

ET F\*S; SF6; S cp; cp; F cp

L54 ANSWER 6 OF 9 HCAPLUS COPYRIGHT 2004 ACS on STN

AN 2000:772535 HCAPLUS

DN 133:312399

ED Entered STN: 03 Nov 2000

TI Cover gas mixtures for protection of reactive **Mg** or **Mg** -alloy **melt**

IN Ricketts, Nigel Jeffrie; Frost, Malcolm Timothy; Cashion, Simon Paul; Korn, Craig John; Baker, Phillip Wilmott

PA Cast Centre Pty Ltd, Australia

SO PCT Int. Appl., 20 pp.

CODEN: PIXXD2

DT Patent

LA English

IC ICM B22D021-04

ICS B22D021-02; A62D001-02; C22B026-22; C22B004-02; C22C001-02; C22C023-00

*applicants*

CC 56-2 (Nonferrous Metals and Alloys)

Section cross-reference(s): 50

FAN.CNT 1

|      | PATENT NO.    | KIND   | DATE     | APPLICATION NO.   | DATE     |
|------|---------------|--|----------|-------------------|----------|
| PI   | WO 2000064614 | A1   | 20001102 | WO 2000-AU393     | 20000428 |
|      | W:            | AE, AG, AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, CA, CH, CN, CR, CU, CZ, DE, DK, DM, DZ, EE, ES, FI, GB, GD, GE, GH, GM, HR, HU, ID, IL, IN, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MA, MD, MG, MK, MN, MW, MX, NO, NZ, PL, PT, RO, RU, SD, SE, SG, SI, SK, SL, TJ, TM, TR, TT, TZ, UA, UG, US, UZ, VN, YU, ZA, ZW, AM, AZ, BY, KG, KZ, MD, RU, TJ, TM |          |                   |          |
|      | RW:           | GH, GM, KE, LS, MW, SD, SL, SZ, TZ, UG, ZW, AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE, BF, BJ, CF, CG, CI, CM, GA, GN, GW, ML, MR, NE, SN, TD, TG   |          |                   |          |
|      | BR 2000010137 | A  | 20020122 | BR 2000-10137     | 20000428 |
|      | TR 200103096  | T2   | 20020422 | TR 2001-200103096 | 20000428 |
|      | EP 1204499    | A1   | 20020515 | EP 2000-920274    | 20000428 |
|      | R:            | AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, MC, PT, IE, SI, LT, LV, FI, RO, MK, CY, AL   |          |                   |          |
|      | JP 2002541999 | T2   | 20021210 | JP 2000-613596    | 20000428 |
|      | AU 766844     | B2   | 20031023 | AU 2000-40930     | 20000428 |
|      | NZ 515084     | A  | 20031031 | NZ 2000-515084    | 20000428 |
|      | TW 500805     | B  | 20020901 | TW 2000-89108130  | 20000504 |
|      | NO 2001005264 | A  | 20011221 | NO 2001-5264      | 20011026 |
|      | ZA 2001008862 | A  | 20021028 | ZA 2001-8862      | 20011026 |
|      | BG 106138     | A  | 20020628 | BG 2001-106138    | 20011123 |
| PRAI | AU 1999-15    | A  | 19990428 |                   |          |
|      | WO 2000-AU393 | W  | 20000428 |                   |          |

## CLASS

| PATENT NO.    | CLASS | PATENT FAMILY CLASSIFICATION CODES                                     |
|---------------|-------|--|
| WO 2000064614 | ICM   | B22D021-04   |
|               | ICS   | B22D021-02; A62D001-02; C22B026-22; C22B004-02; C22C001-02; C22C023-00 |

AB The cover-gas blend for protection of the **Mg** or **Mg** alloy **molten** bath (especially in casting) contains: (a) hydrofluorocarbon type inhibiting gas, typically as difluoromethane or heptafluoropropane at <1% by volume; and (b) suitable carrier gas, especially dry

air, CO<sub>2</sub>, Ar, and/or N<sub>2</sub>. The component gases are selected to have no O<sub>3</sub>-depletion potential, as well as a low Global-Warming potential. The cover-gas blends are suitable for preventing oxidation of **Mg** alloy **melt** for ingot casting, as well as for extinguishing of accidental fires on the molten bath. The protective mixture suitable for preventing surface oxidation on the cast ingots of AZ91 Mg-Al alloy contains 0.04% by volume of 1,1,1,2-tetrafluoroethane in dry air, vs. surface dross and discoloration on the ingots solidified under the similar amount of SF<sub>6</sub> in dry air.

ST **magnesium** alloy **melt** casting cover gas protection;  
hydrofluorocarbon cover gas magnesium fire prevention

IT Hydrocarbons, uses

RL: TEM (Technical or engineered material use); USES (Uses)  
(fluoro, gas mixts. with; cover gas mixts. with hydrofluorocarbons for protection of reactive **Mg-melt** bath)

IT Safety

(foundry, for Mg alloy casting; cover gas mixts. with hydrofluorocarbons for protection of reactive **Mg-alloy melt**)

IT Casting of metals  
 (of Mg-alloys; cover gas mixts. with hydrofluorocarbons for protection  
 of reactive Mg-alloy melt)

IT Fire  
 (prevention, on Mg-alloy melt; cover gas mixts.  
 with hydrofluorocarbons for protection of reactive Mg-alloy  
 melt)

IT 12634-54-7, AZ91  
 RL: PEP (Physical, engineering or chemical process); PROC (Process)  
 (casting of, surface protection in; cover gas mixts. with  
 hydrofluorocarbons for protection of reactive Mg-melt  
 bath)

IT 124-38-9, Carbon dioxide, uses 7440-37-1, Argon, uses 7727-37-9,  
 Nitrogen, uses  
 RL: MOA (Modifier or additive use); USES (Uses)  
 (gas mixts. with; cover gas mixts. with hydrofluorocarbons for  
 protection of reactive Mg-melt bath)

IT 75-10-5, Difluoromethane 75-37-6 354-33-6,  
 Pentafluoroethane 811-97-2, 1,1,1,2-Tetrafluoroethane  
 33660-75-2, Heptafluoropropane 138495-42-8, HFC-43-10  
 163702-05-4, HFE-7200 219484-64-7, HFE-7100  
 RL: TEM (Technical or engineered material use); USES (Uses)  
 (gas mixts. with; cover gas mixts. with hydrofluorocarbons for  
 protection of reactive Mg-melt bath)

IT 7439-95-4, Magnesium, uses  
 RL: TEM (Technical or engineered material use); USES (Uses)  
 (molten; cover gas mixts. with hydrofluorocarbons for  
 protection of reactive Mg-melt bath)

RE.CNT 6 THERE ARE 6 CITED REFERENCES AVAILABLE FOR THIS RECORD

RE

(1) Dougherty; US 5115868 A 1992 HCAPLUS  
 (2) Great Lakes Chemical Corporation; WO 9102564 A 1991  
 (3) Li; US 5855647 A 1999 HCAPLUS  
 (4) Minnesota Mining And Manufacturing Company; WO 9622129 A 1996 HCAPLUS  
 (5) Radzilowski; US 4214899 A 1980 HCAPLUS  
 (6) Tokai Rika Denki Kk; JP 08143985 1996 HCAPLUS

IT 75-10-5, Difluoromethane 75-37-6 354-33-6,  
 Pentafluoroethane 811-97-2, 1,1,1,2-Tetrafluoroethane  
 33660-75-2, Heptafluoropropane 138495-42-8, HFC-43-10  
 163702-05-4, HFE-7200 219484-64-7, HFE-7100  
 RL: TEM (Technical or engineered material use); USES (Uses)  
 (gas mixts. with; cover gas mixts. with hydrofluorocarbons for  
 protection of reactive Mg-melt bath)

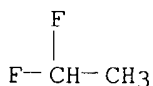
RN 75-10-5 HCAPLUS

CN Methane, difluoro- (8CI, 9CI) (CA INDEX NAME)

F-CH<sub>2</sub>-F

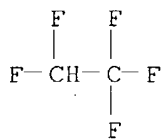
RN 75-37-6 HCAPLUS

CN Ethane, 1,1-difluoro- (6CI, 8CI, 9CI) (CA INDEX NAME)



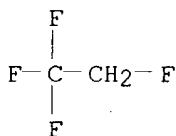
RN 354-33-6 HCAPLUS

CN Ethane, pentafluoro- (6CI, 7CI, 8CI, 9CI) (CA INDEX NAME)



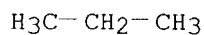
RN 811-97-2 HCAPLUS

CN Ethane, 1,1,1,2-tetrafluoro- (6CI, 7CI, 8CI, 9CI) (CA INDEX NAME)



RN 33660-75-2 HCAPLUS

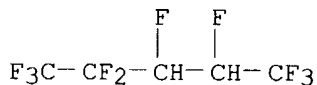
CN Propane, heptafluoro- (6CI, 7CI, 8CI, 9CI) (CA INDEX NAME)



7 (D1-F)

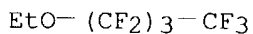
RN 138495-42-8 HCAPLUS

CN Pentane, 1,1,1,2,2,3,4,5,5,5-decafluoro- (9CI) (CA INDEX NAME)



RN 163702-05-4 HCAPLUS

CN Butane, 1-ethoxy-1,1,2,2,3,3,4,4,4-nonafluoro- (9CI) (CA INDEX NAME)



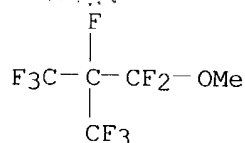
RN 219484-64-7 HCAPLUS

CN Butane, 1,1,1,2,2,3,3,4,4-nonafluoro-4-methoxy-, mixt. with 2-(difluoromethoxymethyl)-1,1,1,2,3,3,3-heptafluoropropane (9CI) (CA INDEX NAME)

CM 1

CRN 163702-08-7

CMF C5 H3 F9 O



CM 2

CRN 163702-07-6  
CMF C5 H3 F9 O

MeO-(CF<sub>2</sub>)<sub>3</sub>-CF<sub>3</sub>

IT 7439-95-4, Magnesium, uses  
RL: TEM (Technical or engineered material use); USES (Uses)  
(molten; cover gas mixts. with hydrofluorocarbons for  
protection of reactive Mg-melt bath)  
RN 7439-95-4 HCAPLUS  
CN Magnesium (8CI, 9CI) (CA INDEX NAME)

Mg

L54 ANSWER 7 OF 9 HCAPLUS COPYRIGHT 2004 ACS on STN  
AN 1999:811058 HCAPLUS  
DN 132:54870  
ED Entered STN: 24 Dec 1999  
TI Pharmaceutical formulations for aerosols with two or more active  
substances  
IN McNamara, Daniel P.; Destefano, George A.  
PA Boehringer Ingelheim Pharmaceuticals, Inc., USA  
SO PCT Int. Appl., 18 pp.  
CODEN: PIXXD2  
DT Patent  
LA English  
IC ICM A61K009-12  
ICS A61K031-135; A61K031-46  
CC 63-6 (Pharmaceuticals)  
FAN.CNT 1

|    | PATENT NO.  | KIND | DATE     | APPLICATION NO.  | DATE     |
|----|---|------|----------|------------------|----------|
| PI | WO 9965464  | A1   | 19991223 | WO 1999-US12785  | 19990608 |
|    | W: AE, AU, BG, BR, BY, CA, CN, CZ, EE, HR, HU, ID, IL, IN, JP, KR, KZ, LT, LV, MX, NO, NZ, PL, RO, RU, SG, SI, SK, TR, UA, UZ, VN, YU, ZA |      |          |                  |          |
|    | RW: AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE  |      |          |                  |          |
|    | DE 19827178   | A1   | 20000427 | DE 1998-19827178 | 19980618 |
|    | DE 19842963   | A1   | 20000323 | DE 1998-19842963 | 19980919 |
|    | CA 2335065  | AA   | 19991223 | CA 1999-2335065  | 19990608 |
|    | AU 9945521  | A1   | 20000105 | AU 1999-45521    | 19990608 |
|    | AU 759222   | B2   | 20030410 |                  |          |

|   |    |          |                   |          |
|---|----|----------|-------------------|----------|
| BR 9911351  | A  | 20010313 | BR 1999-11351     | 19990608 |
| EP 1087750  | A1 | 20010404 | EP 1999-928458    | 19990608 |
| EP 1087750  | B1 | 20031112 |                   |          |
| R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, MC, PT, IE, SI, LT, LV, FI, RO |    |          |                   |          |
| TR 200003721  | T2 | 20010621 | TR 2000-200003721 | 19990608 |
| ZA 200007581  | A  | 20020402 | ZA 2000-7581      | 19990608 |
| EE 200000759  | A  | 20020415 | EE 2000-759       | 19990608 |
| NZ 509418   | A  | 20030630 | NZ 1999-509418    | 19990608 |
| JP 2003522102   | T2 | 20030722 | JP 2000-554344    | 19990608 |
| AT 253896   | E  | 20031115 | AT 1999-928458    | 19990608 |
| RU 2218152  | C2 | 20031210 | RU 2001-101898    | 19990608 |
| PT 1087750  | T  | 20040227 | PT 1999-928458    | 19990608 |
| ES 2211108  | T3 | 20040701 | ES 1999-928458    | 19990608 |
| TW 528606   | B  | 20030421 | TW 1999-88110010  | 19990615 |
| BG 105033   | A  | 20010928 | BG 2000-105033    | 20001207 |
| NO 2000006318   | A  | 20010130 | NO 2000-6318      | 20001212 |
| HR 2000000867   | A1 | 20011031 | HR 2000-867       | 20001215 |
| PRAI DE 1998-19827178   | A  | 19980618 |                   |          |
| DE 1998-19842963  | A  | 19980919 |                   |          |
| WO 1999-US12785   | W  | 19990608 |                   |          |

## CLASS

| PATENT NO. | CLASS   | PATENT FAMILY CLASSIFICATION CODES |
|------------|---|------------------------------------|
| WO 9965464 | ICM   | A61K009-12                         |
|            | ICS   | A61K031-135; A61K031-46            |
| AB         | The present invention relates to new pharmaceutical formulations for aerosols with at least two or more active substances for administration by inhalation or by nasal route. Specifically, the invention relates to pharmaceutical preps. for propellant-driven metered dose aerosols using a fluorohydrocarbon (HFC) as propellant, which contain a combination of active substance of two or more active substances, wherein at least one active substance is present in dissolved form together with at least one other active substance in the form of suspended particles. An aerosol composition contained ipratropium bromide 37 mg, salbutamol sulfate 210.5 mg, citric acid 4 mg, liquefied TG 134a 89.96 g, ethanol 10.03 g, and oleic acid 50 mg. |                                    |
| ST         | fluorohydrocarbon propellant aerosol ipratropium salbutamol   |                                    |
| IT         | Fatty acids, biological studies   |                                    |
|            | RL: THU (Therapeutic use); BIOL (Biological study); USES (Uses)<br>(C5-20; aerosols containing two or more active substances and fluorohydrocarbon propellant)  |                                    |
| IT         | Carbohydrates, biological studies   |                                    |
|            | Glycerides, biological studies  |                                    |
|            | Lecithins   |                                    |
|            | RL: THU (Therapeutic use); BIOL (Biological study); USES (Uses)<br>(aerosols containing two or more active substances and fluorohydrocarbon propellant)   |                                    |
| IT         | Drug delivery systems   |                                    |
|            | (aerosols; aerosols containing two or more active substances and fluorohydrocarbon propellant)  |                                    |
| IT         | Quaternary ammonium compounds, biological studies   |                                    |
|            | RL: THU (Therapeutic use); BIOL (Biological study); USES (Uses)<br>(alkylbenzyltrimethyl, chlorides; aerosols containing two or more active substances and fluorohydrocarbon propellant)  |                                    |
| IT         | Fatty acids, biological studies   |                                    |
|            | RL: THU (Therapeutic use); BIOL (Biological study); USES (Uses)<br>(esters, C5-20; aerosols containing two or more active substances and fluorohydrocarbon propellant)  |                                    |

IT Alcohols, biological studies  
 RL: THU (Therapeutic use); BIOL (Biological study); USES (Uses)  
 (fatty, C5-20; aerosols containing two or more active substances and fluorohydrocarbon propellant)

IT Hydrocarbons, biological studies  
 RL: THU (Therapeutic use); BIOL (Biological study); USES (Uses)  
 (fluoro; aerosols containing two or more active substances and fluorohydrocarbon propellant)

IT Drug delivery systems  
 (inhalants; aerosols containing two or more active substances and fluorohydrocarbon propellant)

IT Drug delivery systems  
 (nasal sprays; aerosols containing two or more active substances and fluorohydrocarbon propellant)

IT 64-17-5, Ethanol, biological studies 112-80-1, Oleic acid, biological studies 431-89-0, TG 227 586-06-1, Orciprenaline **811-97-2**, TG 134a 1338-43-8, Sorbitan monooleate 1944-12-3, Fenoterol hydrobromide 3385-03-3, Flunisolide 4419-39-0, Beclomethasone 12441-09-7D, Sorbitan, esters 13392-18-2, Fenoterol 16110-51-3, Cromoglycic acid 18559-94-9, Salbutamol 22254-24-6, Ipratropium bromide 23031-25-6, Terbutalin 26266-58-0, Sorbitan trioleate 29116-98-1, Sorbitan dioleate 30286-75-0, Oxitropium bromide 51022-70-9, Albuterol sulfate 51333-22-3, Budesonide 54063-54-6, Reproterol 60205-81-4, Ipratropium 69049-73-6, Nedocromil 89365-50-4, Salmeterol 90566-53-3, Fluticasone 148430-28-8  
 RL: THU (Therapeutic use); BIOL (Biological study); USES (Uses)  
 (aerosols containing two or more active substances and fluorohydrocarbon propellant)

IT 50-81-7, Ascorbic acid, biological studies 60-00-4, EDTA, biological studies 77-92-9, Citric acid, biological studies 7647-01-0, Hydrochloric acid, biological studies 7664-38-2, Phosphoric acid, biological studies 7664-93-9, Sulfuric acid, biological studies 7697-37-2, Nitric acid, biological studies  
 RL: THU (Therapeutic use); BIOL (Biological study); USES (Uses)  
 (as stabilizer; aerosols containing two or more active substances and fluorohydrocarbon propellant)

RE.CNT 5 THERE ARE 5 CITED REFERENCES AVAILABLE FOR THIS RECORD

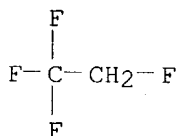
RE

(1) Bell, A; WO 9801147 A 1998 HCAPLUS  
 (2) Henry, R; US 5589156 A 1996 HCAPLUS  
 (3) Jager, P; WO 9413262 A 1994 HCAPLUS  
 (4) McNamara, D; US 5603918 A 1997 HCAPLUS  
 (5) Riker Laboratories Inc; EP 0499344 A 1992 HCAPLUS

IT **811-97-2**, TG 134a  
 RL: THU (Therapeutic use); BIOL (Biological study); USES (Uses)  
 (aerosols containing two or more active substances and fluorohydrocarbon propellant)

RN 811-97-2 HCAPLUS

CN Ethane, 1,1,1,2-tetrafluoro- (6CI, 7CI, 8CI, 9CI) (CA INDEX NAME)



L54 ANSWER 8 OF 9 HCAPLUS COPYRIGHT 2004 ACS on STN  
 AN 1995:422771 HCAPLUS  
 DN 122:165333  
 ED Entered STN: 18 Mar 1995  
 TI Parting compounds containing fluorocarbons which contain no chlorine and may contain hydrogen  
 IN Buchwald, Hans; Hellmann, Joachim; Raschkowski, Boleslaus  
 PA Solvay Fluor und Derivate GmbH, Germany  
 SO Ger. Offen., 6 pp.  
 CODEN: GWXXBX  
 DT Patent  
 LA German  
 IC ICM C10M111-02  
 ICS B29C033-60; C03B040-02  
 ICA C09K015-30; C09K015-22; C09K015-18; C09K003-30; B01F017-00; B22C001-14  
 ICI C10M111-02, C10M105-52, C10M105-34, C10M105-02, C10M105-76, C10M105-24, C10M105-72, C10M107-04, C10M107-18, C10N040-36  
 CC 51-11 (Fossil Fuels, Derivatives, and Related Products)  
 Section cross-reference(s): 38, 57  
 FAN.CNT 1

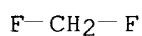
|      | PATENT NO.      | KIND | DATE     | APPLICATION NO. | DATE     |
|------|-----------------|------|----------|-----------------|----------|
| PI   | DE 4321288      | A1   | 19950105 | DE 1993-4321288 | 19930626 |
|      | US 5476603      | A    | 19951219 | US 1994-262930  | 19940621 |
|      | FI 9403069      | A    | 19941227 | FI 1994-3069    | 19940623 |
|      | JP 07150188     | A2   | 19950613 | JP 1994-141905  | 19940623 |
| PRAI | DE 1993-4321288 | A    | 19930626 |                 |          |

## CLASS

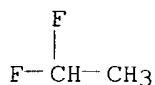
|    | PATENT NO.   | CLASS | PATENT FAMILY CLASSIFICATION CODES   |
|----|--|-------|--|
|    | DE 4321288   | ICM   | C10M111-02   |
|    |  | ICS   | B29C033-60; C03B040-02   |
|    |  | ICA   | C09K015-30; C09K015-22; C09K015-18; C09K003-30; B01F017-00; B22C001-14   |
|    |  | ICI   | C10M111-02, C10M105-52, C10M105-34, C10M105-02, C10M105-76, C10M105-24, C10M105-72, C10M107-04, C10M107-18, C10N040-36 |
|    | US 5476603   | ECLA  | B29C033/60; B29C033/62; B29C033/64   |
| AB | The parting compds. are pressure-liquefied and contain a fluorocarbon chosen from among R134a, R143a, R125, R32, R143, R134, R23, R152a, R14, and R116, a parting agent (e.g., wax, paraffins, silicones, metallic soaps, fats, oils, polymers, or inorg. powders), and optionally solvents (e.g., a gasoline fraction b.p. 100-140°). |       |  |
| ST | parting agent fluorocarbon   |       |  |
| IT | Solvents   |       |  |
|    | (pressure-liquefied fluorocarbon-containing parting compns.)   |       |  |
| IT | Fats and Glyceridic oils   |       |  |
|    | Gasoline   |       |  |
|    | Oils   |       |  |
|    | Paraffin waxes and Hydrocarbon waxes, uses   |       |  |
|    | Polymers, uses   |       |  |
|    | Siloxanes and Silicones, uses  |       |  |
|    | Soaps  |       |  |
|    | Waxes and Waxy substances  |       |  |
|    | RL: TEM (Technical or engineered material use); USES (Uses)  |       |  |
|    | (pressure-liquefied fluorocarbon-containing parting compns.)   |       |  |
| IT | Hydrocarbons, uses   |       |  |
|    | RL: TEM (Technical or engineered material use); USES (Uses)  |       |  |
|    | (fluoro, pressure-liquefied fluorocarbon-containing parting compns.)   |       |  |



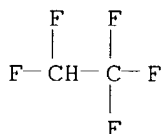
IT 75-10-5, Difluoromethane 75-37-6, R152a 75-46-7, R23  
 75-73-0, R14 76-16-4, R116 354-33-6, R125 359-35-3, R134  
 420-46-2, R143a 430-66-0, R143 557-04-0, Magnesium stearate  
 811-97-2, R134a 7429-90-5D, Aluminum, soaps 7439-92-1D, Lead,  
 soaps 7439-95-4D, Magnesium, soaps 7440-66-6D, Zinc,  
 soaps 7440-70-2D, Calcium, soaps  
 RL: TEM (Technical or engineered material use); USES (Uses)  
 (pressure-liquefied fluorocarbon-containing parting compns.)  
 IT 75-10-5, Difluoromethane 75-37-6, R152a 354-33-6  
 , R125 811-97-2, R134a 7439-95-4D, Magnesium  
 , soaps  
 RL: TEM (Technical or engineered material use); USES (Uses)  
 (pressure-liquefied fluorocarbon-containing parting compns.)  
 RN 75-10-5 HCAPLUS  
 CN Methane, difluoro- (8CI, 9CI) (CA INDEX NAME)



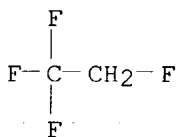
RN 75-37-6 HCAPLUS  
 CN Ethane, 1,1-difluoro- (6CI, 8CI, 9CI) (CA INDEX NAME)



RN 354-33-6 HCAPLUS  
 CN Ethane, pentafluoro- (6CI, 7CI, 8CI, 9CI) (CA INDEX NAME)



RN 811-97-2 HCAPLUS  
 CN Ethane, 1,1,1,2-tetrafluoro- (6CI, 7CI, 8CI, 9CI) (CA INDEX NAME)



RN 7439-95-4 HCAPLUS  
 CN Magnesium (8CI, 9CI) (CA INDEX NAME)

Mg

L54 ANSWER 9 OF 9 HCAPLUS COPYRIGHT 2004 ACS on STN

AN 1954:64093 HCAPLUS

DN 48:64093

OREF 48:11316h-i,11317a-g

ED Entered STN: 22 Apr 2001

TI Pyrolyses of the salts of the perfluoro carboxylic acids

AU La Zerte, J. D.; Hals, L. J.; Reid, T. S.; Smith, G. H.

CS Minnesota Mining &amp; Manufg. Co., St. Paul

SO Journal of the American Chemical Society (1953), 75, 4525-8

CODEN: JACSAT; ISSN: 0002-7863

DT Journal

LA Unavailable

CC 10 (Organic Chemistry)

AB The thermal decomposition of a number of salts of the straight-chain perfluoro acids has been investigated. From the Na salts, terminally unsatd. perfluoroolefins were prepared in yields ranging from 65 to 100%. The reaction is represented by the equation  $C_nF_{2n+1}CF_2CF_2CO_2Na \rightarrow C_nF_{2n+1}CF:CF_2 + CO_2 + NaF$ . Salts of other metals of the groups I, II, and III of the periodic table gave varying yields of olefins.  $C_3F_7CO_2Ag$  (I) and  $C_7F_{15}CO_2Ag$  (II) decomposed to give  $C_6F_{14}$  and  $C_{14}F_{30}$ , resp.  $C_2F_4$  was formed when a mixture of  $CF_3CO_2Na$  and NaOH was heated. A series of fluorocarbon hydrides,  $C_nF_{2n+1}H$ , was prepared by heating the salts of perfluoro acids in  $(CH_2OH)_2$ . The  $NH_4$ , Li, Na, K, Ca, Sr, and Ba salts of the perfluoro acids were all prepared by neutralizing an aqueous solution of

the

acid with a solution of the hydroxide.  $(C_3F_7CO_2)_2Mg$  and  $(C_3F_7CO_2)_2Pb$  were obtained from aqueous  $C_3F_7CO_2H$  (III) and the metal oxides at slightly above  $25^\circ$ ; both salts were hygroscopic; the vacuum-dried Pb salt was further dried by azeotropic distillation with  $CCl_4$ . I and II were prepared by treating freshly prepared  $Ag_2O$  with the dilute aqueous acids.  $(C_3F_7CO_2)_2Cu$  was obtained by passing dry air into a mixture of finely divided Cu powder and excess III at  $120^\circ$ .  $(C_3F_7CO_2)_3Al$  was prepared by the method of Hood and Ihde (C.A. 44, 7228i) from  $AlCl_3$  and excess III in the presence of  $(C_3F_7CO_2)_2O$  (IV) at  $100^\circ$ . The purity of the salts had a great influence on the decomposition reaction. In the presence of an inorg. base, the pyrolysis of the salts gave products contaminated with fluorocarbon monohydrides; to avoid this, the pH of the salt solns. was adjusted to pH 5-7.  $H_2O$  vapors in the pyrolysis zone also led to the formation of H-containing compds. The pyrolyses were carried out, in general, in Pyrex flasks; the rate of the decomposition was controlled by varying the

temperature; the

resulting volatile products were passed through 2 scrubbers containing 15% KOH, dried over  $P_2O_5$ , and condensed in a cold trap. The thermal stabilities of some salts of III were determined by heating small weighed samples 0.5 hr. at  $20-5^\circ$  intervals until almost complete decomposition was obtained; the temperature at which 20% decomposition was obtained (given)

was for

the following salts:  $NH_4$   $185^\circ$ , K  $200^\circ$ , Na  $235^\circ$ , Ba  $275^\circ$ , Sr  $275^\circ$ , Ag  $295^\circ$ ; and for  $(CF_3)_2CFCO_2Na$   $185^\circ$ .

The Na salts of higher straight-chain perfluoro acids

underwent 20% decomposition at  $240-50^\circ$ , and  $C_4F_9CO_2K$  at  $175-80^\circ$ .

The Na and Ba salts of  $CF_3CO_2H$  gave  $CF_3COF$  and  $(CF_3CO)_2O$ ; the same

products were obtained from the Li and Ca salts. The pyrolysis of

$CF_3CO_2Na$  in the presence of solid NaOH proceeded at about  $270^\circ$

exothermically to give  $C_2F_4$ , along with some  $CF_3COF$  and  $CHF_3$ ; the min.

yield of  $C_2F_4$  was 32% in better than 98% purity; 1% by weight of  $Pr_3N$  was

always added to the  $C_2F_4$  to prevent the explosive polymerization of the

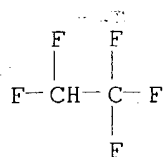
monomer. The following salts of III were pyrolyzed and the pyrolysis

products determined (the decomposition temperature, % yield  $CF_3CF:CF_2$ , and the

other

fluorinated products formed given): Li, 240-50°, 20, C3F7COF (V), IV, III; K, 215-35°, 98, -, Mg, 275-300°, <5, high-boiling liquid; Ca, 275-300°, <10, V, IV, III; Sr, 275-85°, 25, V, III; Ba, 265-75°, 78, -, Pb, 300-5°, <10, V, IV, some III; Cu, trace, V, unidentified product; Al, 250°, <5, V, III, C2F6; NH4, 180-200°, 0, CF3CF2CF2H; Ag, 300-20°, 45, C6F14. C4F9CO2K (1681 g.) and 907 g. (CH2OH)2 heated 5 hrs. at 170-90° gave 1169 g. cold-trap condensate which on fractionation yielded 1017 g. (84%) CF3(CF2)3H, b740 14°,  $\lambda_{\text{maximum}}$  3015 cm.<sup>-1</sup> (C-H). Similarly were prepared from the Na salts of the appropriate perfluoro acids the following hydrides CF3(CF2)nH (VI) (n, % yield, b.p./740 mm. given): 1, 98, -50°; 2, 97, -16°; 4, 80, 46°; 6, 60, 94°, n25D 1.2690. C5F11CO2Na (210 g.), prepared in 93% yield by neutralizing C5F11CO2H with aqueous NaOH, pyrolyzed at about 250° yielded 141 g. (90%) C3F7CF:CF2, b. 28-9.0°, n25D 1.2571,  $\lambda_{\text{maximum}}$  1795 cm.<sup>-1</sup> Similarly were prepared the following olefins from the appropriate Na salts (compound, % yield, b.p., and n15D given): C2F4, 90, -74°, -, CF3CF:CF2, 97, -29°, -, C2F5CF:CF2, 91, 1°, -, C5F11CF:CF2, 86, 81, 1.2782; C7F15CF:CF2, 65, 123°, 1.2868. The infrared absorption spectra of VI with n = 2, 3, 4, and 6 all showed C-H absorption in the range 2940-2990 cm.<sup>-1</sup>.

- IT Spectra  
(of hydrocarbons (fluorinated))
- IT Pyrolysis  
(of perfluoro carboxylic acid salts)
- IT Salts  
(of perfluoro carboxylic acids, pyrolysis of)
- IT Octanoic acid, pentadecafluoro-, silver salt  
Propionic acid, 2,3,3,3-tetrafluoro-2-(trifluoromethyl)-, sodium salt  
(pyrolysis of)
- IT 76-05-1, Acetic acid, trifluoro-  
(and derivs.)
- IT 375-22-4, Butyric acid, heptafluoro-  
(and derivs., pyrolysis of)
- IT 116-14-3, Ethylene, tetrafluoro-  
(formation of, from CF3Co2Na and NaOH)
- IT 75-46-7, Fluoroform  
(formation of, from Na trifluoroacetate)
- IT 76-16-4, Ethane, hexafluoro-  
(formation of, in (C3F7CO2)3Al pyrolysis)
- IT 116-15-4, Propene, hexafluoro-  
(formation of, in pyrolysis of heptafluorobutyric acid salts)
- IT 307-62-0, Tetradecane, triacontafluoro- 354-33-6, Ethane,  
pentafluoro- 354-34-7, Acetyl fluoride, trifluoro- 355-42-0, Hexane,  
tetradecafluoro- 355-63-5, 1-Heptene, tetradecafluoro- 357-26-6,  
1-Butene, octafluoro- 375-17-7, Butane, 1,1,1,2,2,3,3,4,4-nonafluoro-  
375-61-1, Pentane, 1,1,1,2,2,3,3,4,4,5,5-undecafluoro- 375-83-7,  
Heptane, 1,1,1,2,2,3,3,4,4,5,5,6,6,7,7-pentadecafluoro- 376-22-7,  
1-Nonene, octadecafluoro- 376-87-4, 1-Pentene, decafluoro- 2252-84-8,  
Propane, 1,1,1,2,2,3,3-heptafluoro-  
(preparation of)
- IT 336-23-2, Valeric acid, nonafluoro-, potassium salt 2923-26-4, Hexanoic  
acid, undecafluoro-, sodium salt  
(pyrolysis of)
- IT 354-33-6, Ethane, pentafluoro-  
(preparation of)
- RN 354-33-6 HCAPLUS
- CN Ethane, pentafluoro- (6CI, 7CI, 8CI, 9CI) (CA INDEX NAME)



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